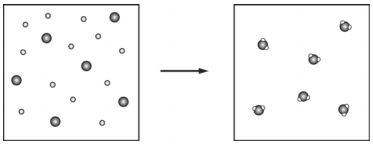
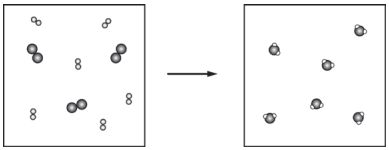
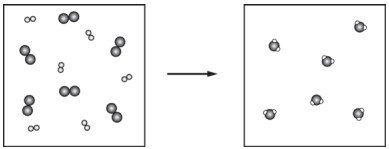
**Honors Chemistry II** Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

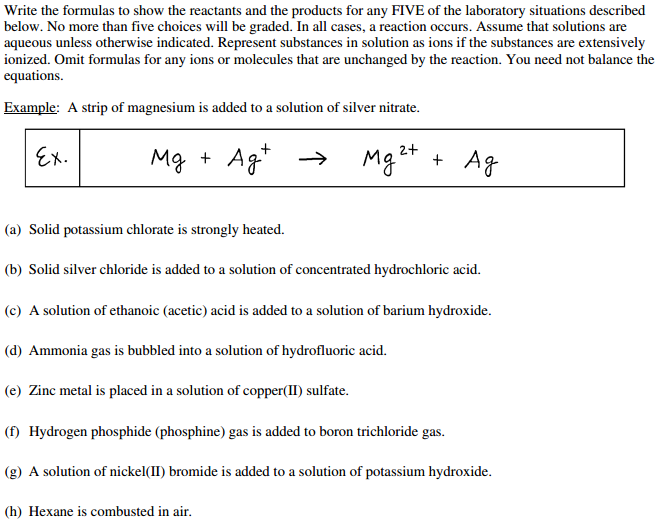
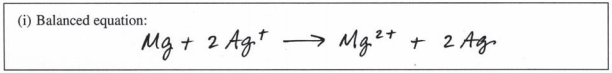
Unit 6 Review

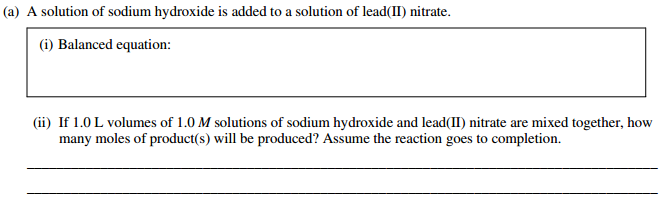
Things you need to have a thorough understanding of (all topics covered in Honors Chemistry and in this class so far are fair game):

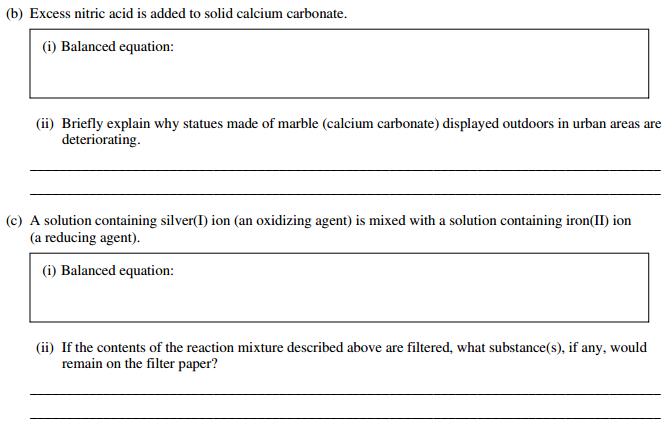
* Naming Compounds
* Balancing Chemical Equations
* Types of Chemical Equations
* Types of Chemical Reactions
* Predicting Products based on stability and types
* Chemical Reactivity
* Reaction Types
  + Acid-base (Arrhenius, Bronsted-Lowry, Lewis)
  + Coordination Complex
  + Precipitation
  + Redox Reactions (simple), most are single replacements

1. Which of the following particulate diagrams best shows the formation of water vapor from hydrogen gas and oxygen gas in a rigid container at 125°C?
   1. 
   2. 
   3. 
   4. 
2. What remains in the reaction vessel after equal masses of K(s) and Cl2(g) have reacted until either one or both of the reactants have been completely consumed?
   1. KCl only
   2. KCl and K only
   3. KCl and Cl2 only
   4. KCl, K, and Cl2

Write balanced net ionic equations for the following reactions.







Key:

